



Cambridge International Examinations
Cambridge International General Certificate of Secondary Education

CO-ORDINATED SCIENCES

0654/11

Paper 1 Multiple Choice

May/June 2015

45 minutes

Additional Materials: Multiple Choice Answer Sheet
Soft clean eraser
Soft pencil (type B or HB is recommended)



READ THESE INSTRUCTIONS FIRST

Write in soft pencil.

Do not use staples, paper clips, glue or correction fluid.

Write your name, Centre number and candidate number on the Answer Sheet in the spaces provided unless this has been done for you.

DO NOT WRITE IN ANY BARCODES.

There are **forty** questions on this paper. Answer **all** questions. For each question there are four possible answers **A, B, C** and **D**.

Choose the **one** you consider correct and record your choice in **soft pencil** on the separate Answer Sheet.

Read the instructions on the Answer Sheet very carefully.

Each correct answer will score one mark. A mark will not be deducted for a wrong answer.

Any rough working should be done in this booklet.

A copy of the Periodic Table is printed on page 20.

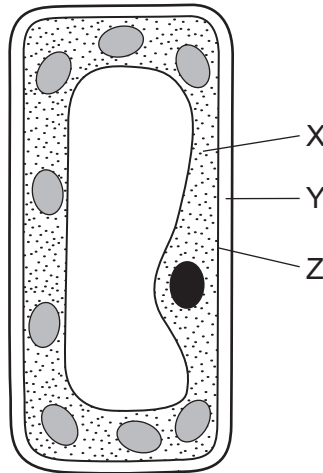
Electronic calculators may be used.

This document consists of **19** printed pages and **1** blank page.

1 Which is a characteristic of all living things?

- A a heart
- B breathing
- C excretion
- D sexual reproduction

2 The diagram shows a typical plant cell.



Which row is correct?

	cell membrane	cell wall	cytoplasm
A	X	Y	Z
B	X	Z	Y
C	Z	X	Y
D	Z	Y	X

3 What is diffusion?

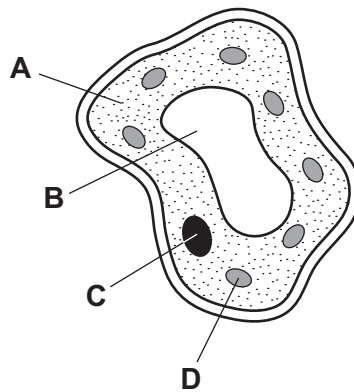
- A the net movement of molecules from a region of their higher concentration to a region of their lower concentration down a concentration gradient
- B the net movement of molecules from a region of their higher concentration to a region of their lower concentration up a concentration gradient
- C the net movement of molecules from a region of their lower concentration to a region of their higher concentration down a concentration gradient
- D the net movement of molecules from a region of their lower concentration to a region of their higher concentration up a concentration gradient

4 What is an enzyme?

- A a carbohydrate that assists in the digestion of the contents of the stomach
- B a chemical that absorbs light for photosynthesis
- C a chemical that alters the activity of a target organ
- D a protein that alters the rate of a chemical reaction

5 The diagram shows a section through a cell from a leaf.

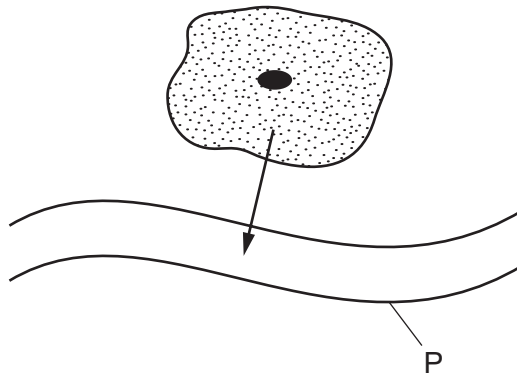
Which part makes simple sugars using light?



6 In a balanced diet, which constituents provide most energy?

- A carbohydrate and protein
- B fat and carbohydrate
- C fat and fibre
- D vitamins and protein

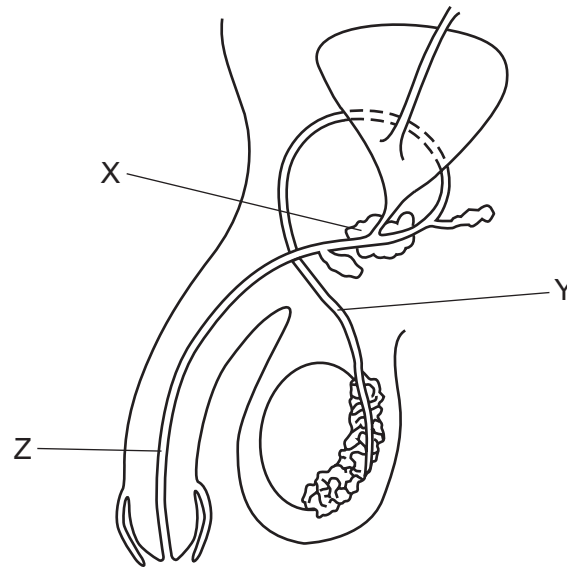
- 7 The arrow shows urea leaving a cell and passing into structure P.



What is P?

- A a capillary
 - B an artery
 - C a vein
 - D the small intestine
- 8 A person touches a hot object which triggers a reflex action.
- In which order does the signal travel in the reflex arc?
- A relay neurone → spinal cord → sensory neurone
 - B sensory neurone → spinal cord → motor neurone
 - C spinal cord → sensory neurone → stimulus
 - D stimulus → motor neurone → spinal cord

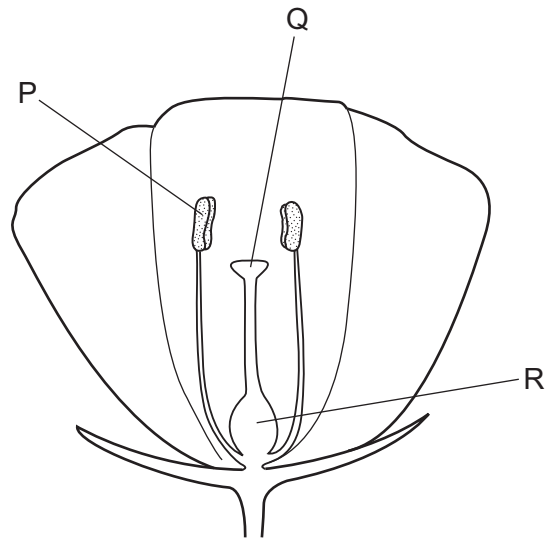
9 The diagram shows the male reproductive system.



Which row identifies structures X, Y and Z?

	urethra	sperm duct	prostate gland
A	X	Y	Z
B	X	Z	Y
C	Z	X	Y
D	Z	Y	X

10 The diagram shows a section through an insect-pollinated flower.



What are the functions of P, Q and R?

	P	Q	R
A	to produce ovules	to produce pollen	to receive pollen
B	to produce pollen	to produce ovules	to receive pollen
C	to produce pollen	to receive pollen	to produce ovules
D	to receive pollen	to produce pollen	to produce ovules

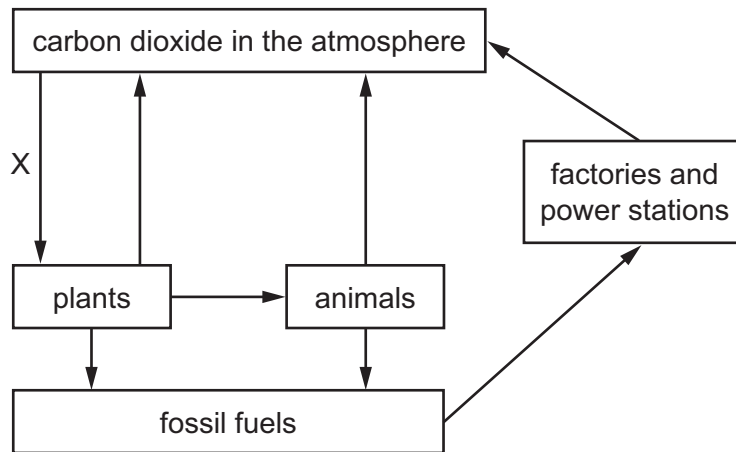
11 Which process is responsible for the flow of energy along a food chain?

- A** excretion
- B** feeding
- C** respiration
- D** seed dispersal

12 Which gas has the biggest greenhouse effect?

- A** carbon monoxide
- B** methane
- C** nitrogen
- D** oxygen

13 The diagram shows part of the carbon cycle.

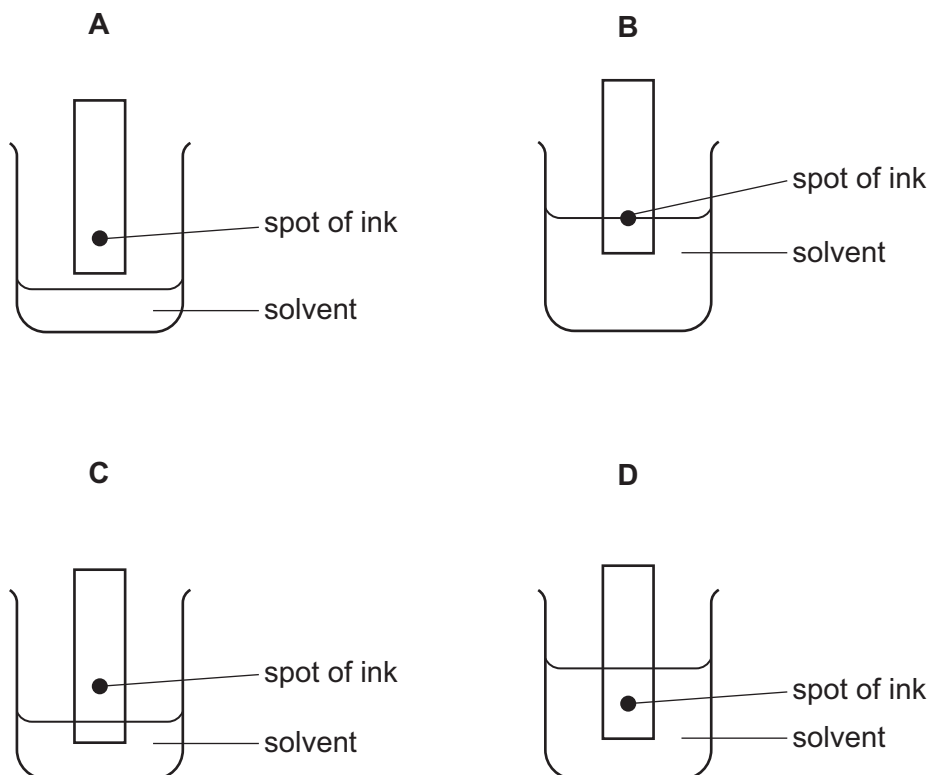


What process does X represent?

- A combustion
- B decay
- C photosynthesis
- D respiration

14 The colours in an ink can be separated by chromatography.

Which diagram shows the correct way to set up the apparatus?



19 Hydrogen and oxygen react explosively to form water.

Which words describe this reaction?

	combustion	oxidation
A	✓	✓
B	✓	x
C	x	✓
D	x	x

key

✓ = yes

x = no

20 Four substances are added to an acid.

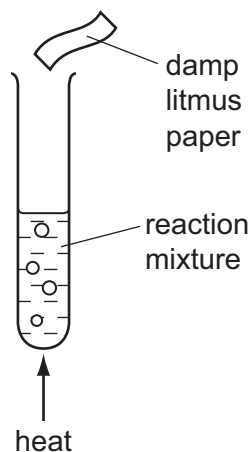
The substances are

- 1 calcium oxide
- 2 magnesium carbonate
- 3 sodium chloride
- 4 sodium hydroxide

Which substances neutralise the acid?

- A** 1 only **B** 1, 2 and 4 **C** 3 and 4 **D** 4 only

21 The diagram shows a chemical reaction that produces a gas.



The gas bleaches damp litmus paper.

What is the gas?

- A ammonia
 - B chlorine
 - C hydrogen
 - D oxygen
- 22 Which statement about the elements in Group VII of the Periodic Table is correct?
- A Chlorine displaces bromine from potassium bromide.
 - B The colour of the elements becomes darker up the group.
 - C The melting point of the elements decreases down the group.
 - D The reactivity of the elements increases down the group.

23 The table shows information about some minerals.

mineral	chemical formula
bauxite	Al_2O_3
galena	PbS
hematite	Fe_2O_3
rutile	TiO_2

Which minerals contain a transition element?

- A bauxite and galena
- B bauxite and hematite
- C galena and rutile
- D hematite and rutile

24 Two tests are done on material Y.

The tests show that Y conducts electricity and is soft.

What is Y?

- A copper
- B lithium
- C sodium chloride
- D sulfur

25 Which process does **not** produce carbon dioxide?

- A complete combustion of fossil fuels
- B reaction of an acid with a carbonate
- C respiration in plants
- D rusting iron

26 Lime is manufactured from limestone and is used for treating industrial waste.

Which row describes the type of reaction involved in the manufacture of lime and in the treatment of industrial waste?

	manufacture	waste treatment
A	reduction	neutralisation
B	reduction	oxidation
C	thermal decomposition	neutralisation
D	thermal decomposition	oxidation

27 A fuel used for cooking food is the hydrocarbon ...1... that burns in an ...2... reaction.

Which words correctly complete gaps 1 and 2?

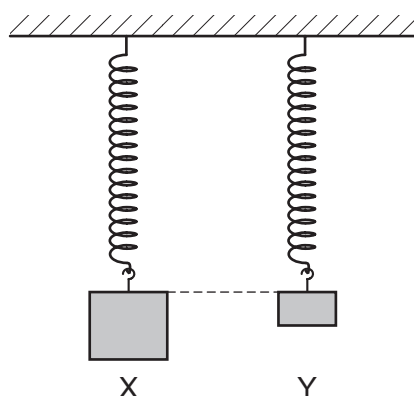
	1	2
A	coke	endothermic
B	coke	exothermic
C	methane	endothermic
D	methane	exothermic

28 The circuit of a motor racing track is 3.0 km in length. In a race, a car goes 25 times round the circuit in 30 minutes.

What is the average speed of the car?

- A** 75 km/hour
- B** 90 km/hour
- C** 150 km/hour
- D** 750 km/hour

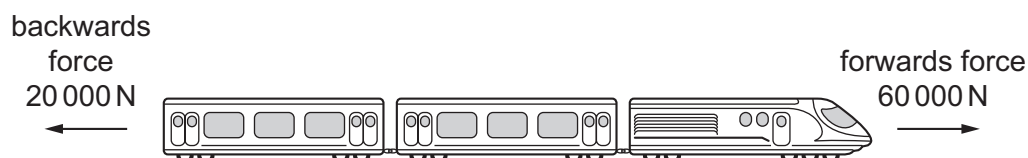
- 29 Two objects X and Y are suspended from identical springs. Both springs extend by the same amount.



What does this show about the masses and about the weights of objects X and Y?

	masses	weights
A	mass X is greater than mass Y	weight X is greater than weight Y
B	mass X is greater than mass Y	weight X is equal to weight Y
C	mass X is equal to mass Y	weight X is equal to weight Y
D	mass X is equal to mass Y	weight X is less than weight Y

- 30 A train travels along a horizontal track at constant speed. Two of the forces acting on the train are shown in the diagram.

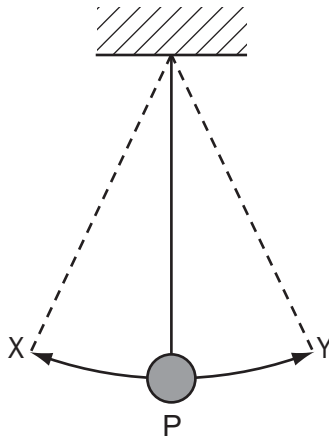


A force of air resistance is also acting on the train to give it a resultant force of zero.

What is this air resistance force?

- A** 40 000 N backwards
- B** 80 000 N backwards
- C** 40 000 N forwards
- D** 80 000 N forwards

- 31 The diagram shows an object attached to a thread, swinging between point X and point Y, passing through point P.



Which row best describes the kinetic energy and the gravitational energy of the object when it is passing through point P?

	kinetic energy	gravitational energy
A	maximum	maximum
B	maximum	minimum
C	minimum	maximum
D	minimum	minimum

- 32 To keep a bottle of fruit juice cool on a hot day, it is wrapped in a cloth soaked in water.

Why is this method successful?

- A** Water has a high boiling point.
 - B** Water has a low melting point.
 - C** Water is a poor conductor of heat.
 - D** Water produces a cooling effect as it evaporates.
- 33 There is a vacuum between the double walls of a vacuum flask.

Which types of heat transfer are reduced by the vacuum?

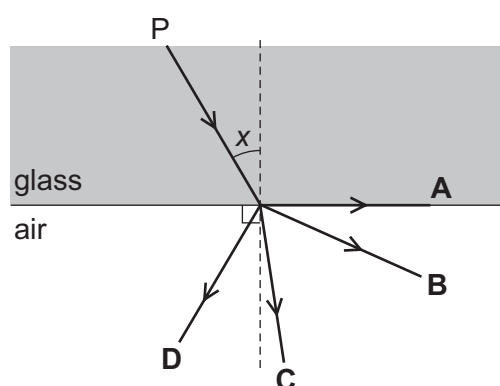
- A** conduction, convection and radiation
- B** conduction and convection only
- C** conduction and radiation only
- D** convection and radiation only

34 Which row gives an example of a longitudinal wave and describes the direction of the vibrations?

	example of a longitudinal wave	vibrations
A	light wave	at right angles to the direction the wave travels
B	light wave	in the same direction as the wave travels
C	sound wave	at right angles to the direction the wave travels
D	sound wave	in the same direction as the wave travels

35 The diagram shows a ray of light travelling from P. Angle x is less than the critical angle.

In which labelled direction does the ray continue?

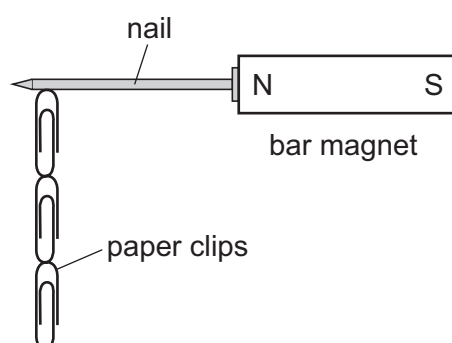


36 Which type of wave **cannot** travel through a vacuum?

- A** infra-red radiation
- B** microwaves
- C** sound waves
- D** X-rays

37 Four nails **A**, **B**, **C** and **D** are tested to find which makes the strongest permanent magnet.

One of the nails is placed against a bar magnet and the number of paper clips which the nail can support is recorded.

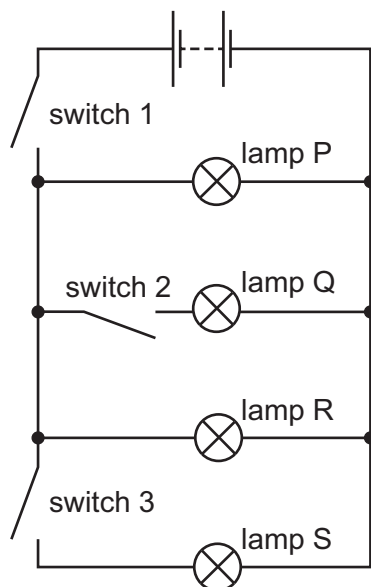


The bar magnet is then removed and the number of paper clips remaining attached to the nail is recorded. Each nail is tested in turn.

Which nail becomes the strongest permanent magnet?

nail	number of paper clips attached to the nail	
	bar magnet present	bar magnet removed
A	2	0
B	2	1
C	4	3
D	5	2

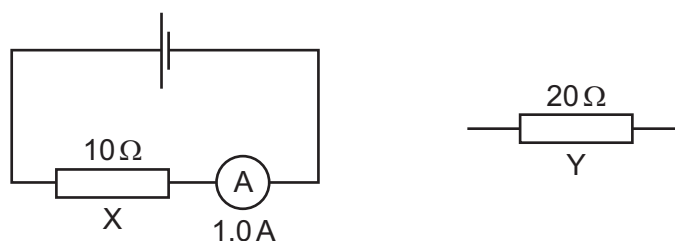
- 38 The circuit shown contains three switches and four lamps P, Q, R and S.



Which switches must be closed to light only lamps P and R?

- A switch 1 only
 - B switch 1 and switch 2 only
 - C switch 1 and switch 3 only
 - D switch 2 and switch 3 only
- 39 The diagram shows a circuit containing a $10\ \Omega$ resistor X and an ammeter. The ammeter reading is 1.0 A .

A $20\ \Omega$ resistor Y is also available.



Which change to the circuit produces a reading on the ammeter that is greater than 1.0 A ?

- A connecting Y in parallel with X
- B placing X on the other side of the ammeter
- C replacing X with Y
- D reversing the connections to X

40 Which type of radiation has the greatest ionising effect?

- A infra-red rays
- B α -particles
- C β -particles
- D γ -rays

DATA SHEET
The Periodic Table of the Elements

		Group											
		I	II	III	IV	V	VI	VII	VIII	IX	X	XI	
		1 H Hydrogen 1											
7	9	20 He Helium 2											
3	4	7	8	9	10	11	12	13	14	15	16	17	18
Li Lithium	Be Beryllium	B Boron	C Carbon	N Nitrogen	O Oxygen	F Fluorine	Ne Neon	Na Sodium	Mg Magnesium	Al Aluminium	Si Silicon	P Phosphorus	S Sulfur
11	12	13	14	15	16	17	18	19	20	21	22	23	24
K Potassium	Ca Calcium	Sc Scandium	Ti Titanium	V Vanadium	Cr Chromium	Mn Manganese	Fe Iron	Co Cobalt	Ni Nickel	Cu Copper	Zn Zinc	Ga Gallium	Ge Germanium
37	38	39	40	41	42	43	44	45	46	47	48	49	50
Rb Rubidium	Sr Strontium	Y Yttrium	Zr Zirconium	Nb Niobium	Mo Molybdenum	Tc Technetium	Ru Ruthenium	Rh Rhodium	Pd Palladium	Ag Silver	Cd Cadmium	In Indium	Sn Tin
55	56	57	58	59	60	61	62	63	64	65	66	67	68
Cs Caesium	Ba Barium	La Lanthanum	Hf Hafnium	Ta Tantalum	W Tungsten	Re Rhenium	Os Osmium	Ir Iridium	Pt Platinum	Au Gold	Hg Mercury	Tl Thallium	Pb Lead
87	88	89	90	91	92	93	94	95	96	97	98	99	100
Fr Francium	Ra Radium	Ac Actinium	Th Thorium	Pa Protactinium	U Uranium	Np Neptunium	Pu Plutonium	Am Americium	Cm Curium	Bk Berkelium	Cf Californium	Es Einsteinium	Fm Fermium
133	137	139	140	141	142	143	144	145	146	147	148	149	150
At Astatine	Po Polonium	Bi Bismuth	Pb Lead	Tl Thallium	Pb Lead	Bi Bismuth	Po Polonium	At Astatine	Rn Radon	Fr Francium	Ra Radium	Ac Actinium	Th Thorium
151	152	153	154	155	156	157	158	159	160	161	162	163	164
Lu Lutetium	Hf Hafnium	Ta Tantalum	W Tungsten	Re Rhenium	Os Osmium	Ir Iridium	Pt Platinum	Au Gold	Hg Mercury	Tl Thallium	Pb Lead	Bi Bismuth	Po Polonium
71	72	73	74	75	76	77	78	79	80	81	82	83	84
Yb Ytterbium	Lu Lutetium	Hf Hafnium	Ta Tantalum	W Tungsten	Re Rhenium	Os Osmium	Ir Iridium	Pt Platinum	Au Gold	Hg Mercury	Tl Thallium	Pb Lead	Bi Bismuth
101	102	103	104	105	106	107	108	109	110	111	112	113	114
La Lanthanum	Ce Cerium	Pr Praseodymium	Nd Neodymium	Pm Promethium	Sm Samarium	Eu Europium	Gd Gadolinium	Tb Terbium	Dy Dysprosium	Ho Holmium	Er Erbium	Tm Thulium	Yb Ytterbium
58-71	90	91	92	93	94	95	96	97	98	99	100	101	102
La Lanthanum	Ce Cerium	Pr Praseodymium	Nd Neodymium	Pm Promethium	Sm Samarium	Eu Europium	Gd Gadolinium	Tb Terbium	Dy Dysprosium	Ho Holmium	Er Erbium	Tm Thulium	Yb Ytterbium
89	90	91	92	93	94	95	96	97	98	99	100	101	102
Ac Actinium	Th Thorium	Pa Protactinium	U Uranium	Np Neptunium	Pu Plutonium	Am Americium	Cm Curium	Bk Berkelium	Cf Californium	Es Einsteinium	Fm Fermium	Md Mendelevium	No Nobelium
87	88	89	90	91	92	93	94	95	96	97	98	99	100
Fr Francium	Ra Radium	Ac Actinium	Th Thorium	Pa Protactinium	U Uranium	Np Neptunium	Pu Plutonium	Am Americium	Cm Curium	Bk Berkelium	Cf Californium	Es Einsteinium	Fm Fermium

*58-71 Lanthanoid series
†90-103 Actinoid series

Key

a	X
b	

a = relative atomic mass
X = atomic symbol
b = proton (atomic) number

The volume of one mole of any gas is 24 dm³ at room temperature and pressure (r.t.p.).